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Unit 9 Solutions Chemistry Packet

1. Calculate the molarity of each of the following solutions: (a) 2.5 mol of NaOH in 0.500 L of solution (b) 1.8L of solution containing 3.3 mol KNO₃ (c) 3.0 mol of NaOH in 0.500 L of solution (d) 12 g of HCl in 250 mL of solution 2. Calculate the total moles of solute in each of the following solutions:

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(a) 2.5 mol of NaOH in 0.500 L of solution (b) 1.8L of solution containing 3.3 mol KNO₃ (c) 30. g of NaOH in 0.500 L of solution (d) 522 g of K₂SO₄ in 1.5 L of solution (e) 12 g of HCl in 250 mL of solution 2. Calculate the total moles of solute in each of the following solutions: 2 Calculate the concentration of various solutions

Practice Packet Unit 9: Solutions

(1) an element and a mixture (3) a solution and a mixture (2) an element and a compound (4) a solution and a compound 12. Two different samples decompose when heated. Only one of the samples is soluble in water. Based on this information, these two samples are (1) both the same element (3) both the same compound

Practice Packet Unit 9: Solutions

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Mr. Dolgos Regents Chemistry NOTE PACKET Unit 9: Acids ...

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Name: _____ Date: _____ Unit 9: Solutions Review Molarity Calculations 1) 2.3 moles of sodium chloride in 0.45 liters of solution. 2) 1.2 moles of calcium carbonate in 1.22 liters of solution. 3) 98 grams of NaOH in 2.2 liters of solution. 4) 1.2 grams HCl in 25 mL of solution.

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Note Packet UNIT 7: SOLUTIONS Unit Vocabulary: Alloy Aqueous boiling point boiling point elevation colligative properties colloid concentrated concentration dilute freezing point depression heterogeneous homogeneous insoluble miscible mixture molarity (M) parts per million (ppm) percent by mass percent by volume precipitate saturated solubility

KEY *KEY* Note Packet UNIT 7: SOLUTIONS - Pace Chemistry

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Name Date Pd Chemistry Unit 9 Solutions & Stoich. Review

The flipped classroom is not an online course. It is a way to engage students and make them more active in the learning process. Students watch a short video lecture for homework (click on the video below for an example) and take notes and complete practice examples.In the classroom these problems are evaluated by the students and teacher to address difficulties and misconceptions.

Mr. Palermo's Flipped Chemistry Classroom - www.mrpalermo.com

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